

Hydrogen incorporation in enstatite in the system MgO–SiO₂–H₂O–NaCl

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Abstract The incorporation of hydrogen in enstatite in a hydrous system containing various amounts of NaCl was investigated at 25 kbar. The hydrogen content in enstatite shows a clear negative correlation to the NaCl-concentration in the system. The most favourable explanation is the reduction of water fugacity due to dilution. Other reasons for the limited hydrogen incorporation at high NaCl levels, such as a significant influence of Na⁺ on the defect chemistry or an exchange between OH[−] and Cl[−] in enstatite, appear much less important. A partition coefficient $D_{\text{Na}}^{\text{En/Fluid}} = 0.0013$ could be determined, demonstrating that Na is less incompatible in enstatite than H. The new results support the idea that dissolved components have to be considered when the total hydrogen storage capacity in nominally anhydrous minerals is estimated, especially in geological settings with high levels of halogens, such as subduction zones.

Keywords Enstatite · Hydrogen incorporation · Water activity · Sodium · Chlorine

Introduction

Significant traces of hydrogen are incorporated into nominally anhydrous minerals of the Earth's upper mantle

(Kitamura et al. 1987; Bell and Rossman 1992; Ingrin and Skogby 2000; Grant et al. 2007), influencing many physical properties (e.g. Mackwell et al. 1985; Karato 1990; Mei and Kohlstedt 2000; Wang et al. 2006; Yoshino et al. 2006) and mantle processes, such as partial melting. More than half of this hydrogen inventory is hosted in orthopyroxene (Grant et al. 2007), which exhibits a rather narrow range around 200 ppm (wt) H₂O in mantle xenoliths from continental alkali basalts and kimberlites (Grant et al. 2007). In contrast, natural mantle orthopyroxenes from sub-arc settings show a much higher variability (40–260 ppm H₂O) and generally lower water contents (Peslier et al. 2002). The incorporation of water in orthopyroxene has also been investigated in many experimental studies (Rauch and Keppler 2002; Stalder 2004; Stalder et al. 2005; Mierdel et al. 2007), where up to several thousands ppm H₂O have been reported. In none of the previous studies the effect of dissolved salts in the coexisting fluids has been examined. NaCl is an important constituent of seawater and allows substantial Cl-incorporation in the basaltic oceanic crust and during serpentinization of mantle rocks at mid ocean ridges. Upon serpentine breakdown the incorporated chlorine is subsequently released into fluid phases during subduction processes (Scambelluri et al. 1997, 2004) and may affect the water activity in the mantle wedge. In this study the effect of NaCl on the hydrogen incorporation in pure enstatite is investigated. As a by-product we present an upper value for Cl-solubility in enstatite, and discuss consequences for the global chlorine cycle.

Experimental and analytical procedure

Orthopyroxene crystals were synthesised applying a method similar to Stalder (2002). The starting mixtures consisted of

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an enstatite + quartz mixture (3:1 by weight), to which solid NaCl or—in one case—KCl was added in different amounts (Table 1). For most syntheses 50–55 mg oxide mixtures and 10–15 mg H₂O were sealed in a Pt-capsule with an outer (inner) diameter of 4.0 (3.6) mm. In the most NaCl-rich synthesis run RS122 only 4.5 mg H₂O was added. All syntheses were performed in an end-loaded piston cylinder apparatus at 25 kbar and temperatures decreasing from 1,400 to 1,150°C with a rate of 10–11°/h. When the final temperature was reached, the run was terminated by switching off the power.

After the run all charges were weighed, pierced and the pH of the extruding fluid was checked with indicator paper. The checked fluid was in all cases neutral to slightly alkaline (pH = 7–9). Subsequently, the pierced capsules were dried in an oven at 110–120°C, whereupon salt crystals formed at the opening of the capsule. Run products of the dried charges consisted of large orthopyroxene crystals (up to 1 mm³), amorphous silicate quench material and salt crystals.

From each experimental charge 2 of the largest crystals were handpicked and orientated by their optical properties parallel (100) and (010) in a thermoplastic resin, and ground and polished (Table 1). After preparation the dimensions of the cuboids were between 150 × 230 × 900 µm and 390 × 480 × 2,000 µm. In two cases (i.e. run RS118 and RS122) the size of the synthesized crystals turned out to be too small to allow alignment of individual crystals parallel (100) and (010). Therefore two wafers—one parallel (100) and one parallel (010)—were prepared. The thickness of these wafers was between 70 and 180 µm.

Polarised IR-spectra were recorded on each polished crystal face parallel to the main refractive indices n_α , n_β and n_γ in transmitted light. All IR measurements were performed with a Perkin Elmer 2000 FTIR-spectrometer coupled to an IR microscope. Measurements were taken using a MIR global light source, a KBr beamsplitter, a MCT-detector and a wire-grid polarizer. For each spectrum 100–300 scans in the 1,000 and 5,000 cm⁻¹ range were acquired with a spectral resolution of 4 cm⁻¹. In contrast to Stalder (2004) and Stalder et al. (2005) background corrections were carried out by adjusting a polynomial baseline at 2,750–2,980, 3,250–3,300 and 3,450–3,750 cm⁻¹. In this way, the broad absorbance feature centered at approximately 3,250 cm⁻¹ was excluded for quantification of water, leading to lower water contents and spectra comparable to Grant et al. (2006). H-contents (Table 1) were quantified by adding the α -, β - and γ -components of the background corrected polarised measurements using the calibration of Libowitzky and Rossman (1997).

After IR measurements all crystals were embedded in epoxy and were analysed with a JEOL JXA 8900 RL

Table 1 Chemical compositions of the studied samples (standard deviation 1 σ in brackets)

Sample	Doped salt	In the system		NaCl or KCl mole fraction ^a	Wt % silica in the fluid ^b	ppm Cl in En		ppm Na in En		Na/10 ⁶ Si		Na/10 ⁶ Si		ppm H ₂ O		H/10 ⁶ Si	
		Wt% H ₂ O	Wt% Na			EMP	LA-ICP-MS	EMP	LA-ICP-MS	EMP	LA-ICP-MS	EMP	LA-ICP-MS	FTIR	FTIR	FTIR	FTIR
En (pure)	–	20.0		0.00	40	22 (15)								170		1,889	
#246	NaCl	21.2	0.3	0.01	25	32 (6)								152		1,689	
#247	NaCl	21.2	0.3	0.01	25	29 (7)								137		1,522	
#243	NaCl	20.1	1.0	0.04	18	71 (16)								125		1,389	
#244	NaCl	20.1	1.0	0.04	18	38 (15)								124		1,378	
#236	NaCl	19.8	2.9	0.10	21	26 (16)								122		1,356	
#237	NaCl	19.8	2.9	0.10	34	34 (10)								109		1,211	
RS118	KCl	19.7		0.14	34	46 (11)								118		1,311	
#238	NaCl	20.3	5.3	0.17	38	26 (11)								106		1,178	
#239	NaCl	20.3	5.3	0.17	38	32 (4)								99		1,100	
RS122	NaCl	7.4	18.4	0.66	2	32 (19)								50		556	

^a Calculated from starting mixtures as $\text{ACl}/(\text{ACl} + \text{H}_2\text{O})$ mole fraction. A = Na or K

^b Estimated after Manning (1994) and Newton and Manning (2000)

electron microprobe using an accelerating voltage of 15 kV and a beam current of 30 nA. Albite, olivine and NaCl were used as standards. For Si and Mg the analysis time lasted 15 s on the peak and 5 s on the background, for Na and Cl the analysis time lasted 150 s on the peak and 75 s on the background on K_{α} emission lines. Five to ten points were measured on each crystal. All analysed crystals were stoichiometric within analytical error. Results for Na and Cl are given in Table 1. Detection limits calculated on the basis of counting statistics of the background signals were 30 ppm for Na and 10 ppm for Cl.

In order to get more information about the alkali content of the alkali-poor crystals, samples were also analysed by LA-ICP-MS. ^{23}Na , ^{29}Si , ^{35}Cl , ^{37}Cl and ^{39}K were measured using a Perkin Elmer DRC II (Sciex, Canada) equipped with a GeoLas Laserablation system (Mikrolas, Göttingen) and a 193 nm Excimer Laser (Lambda Physik, Göttingen). Ar-gas was used to transport laser ablation aerosols into the ICP-MS. Results are given in Table 1.

Results

IR-spectra show for all samples absorption bands at the same position (3,360 and 3,070 cm^{-1}), but with different intensity. The positions of the observed absorption bands are identical to the positions in pure (Na-free) enstatite (Fig. 1). The absorbance in different directions follows the order $E//n_{\gamma} \gg E//n_{\alpha} \geq E//n_{\beta}$. In all samples, the γ -component ($E//n_{\gamma}$) of OH-dipole vectors accounts for 60–70% of the total absorption.

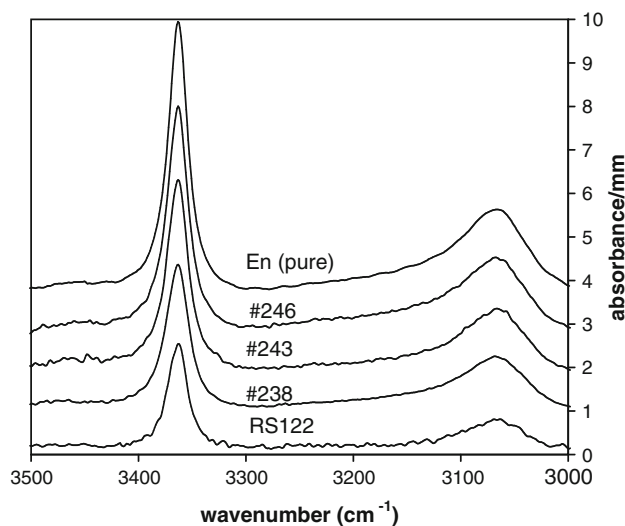


Fig. 1 FTIR-spectra of selected samples synthesized in a broad salinity range (see Table 1). All spectra shown represent the total absorbance of the OH-dipole determined by adding the intensities of the polarised measurements $E//n_{\alpha} + E//n_{\beta} + E//n_{\gamma}$

Na-contents measured by laser ablation-ICP-MS are in quite good agreement with electron microprobe data. In contrast, Cl-contents measured by laser ablation-ICP-MS deviate by approximately 1 order of magnitude from electron microprobe data. The Cl level determined by electron microprobe is just above the detection limit and no correlation between Cl-content of the start mixture and Cl-content in orthopyroxene could be revealed (Table 1). It has to be noted that even electron microprobe analysis may overestimate the real Cl-concentration of the samples, as some Cl may have been excited from the epoxy matrix due to secondary fluorescence. Therefore the values measured by electron microprobe should be regarded as upper limit for the real Cl-concentration. This idea is supported by the fact that even in the pure enstatite sample—synthesized in a nominally Cl-free system—a similar level of Cl is revealed (Table 1). Laser ablation-ICP-MS is not the method of choice to analyse low-level concentrations of chlorine, but the exact reason for the elevated measured Cl-concentrations could not be revealed. Possible explanations are (1) volatilisation from the epoxy during ablation and/or (2) matrix effects. As the Cl-content of nearly all samples is within analytical error constant and independent on the Na-content, a continuous low-level leakage from fluid inclusions as Cl-source can be excluded. Though, laser ablation turned out to be a valuable tool for detecting NaCl-rich inclusions. Solid NaCl inclusions in the synthetic crystals have not been observed under the microscope, but can be inferred from the laser ablation-ICP-MS signal (Fig. 2). In nearly all samples the time resolved ablation signal shows a homogeneous NaCl/Si-ratio (Fig. 2a). Only in the most NaCl-rich run, a simultaneous Na- and Cl-increase at constant Si-signal could be observed (Fig. 2b), suggesting ablation of NaCl-inclusions or NaCl-rich brines in the enstatite crystal.

K was for all laser ablation-ICP-MS analyses below the detection limit (<10 wt. ppm).

Discussion

Reduction of H-incorporation

There are several possible factors to explain the reduced H incorporation with increasing salt content in the system: (1) OH^- is replaced by Cl^- , (2) Na^+ affects the defect chemistry, and (3) the water fugacity is reduced due to the dissolved salt.

Cl-concentrations in enstatite determined by electron microprobe do not exhibit a significant trend over a broad range of NaCl in the system (Table 1). Therefore, a replacement of OH^- by Cl^- cannot account for the strong decrease in H-content with increasing NaCl in the system.

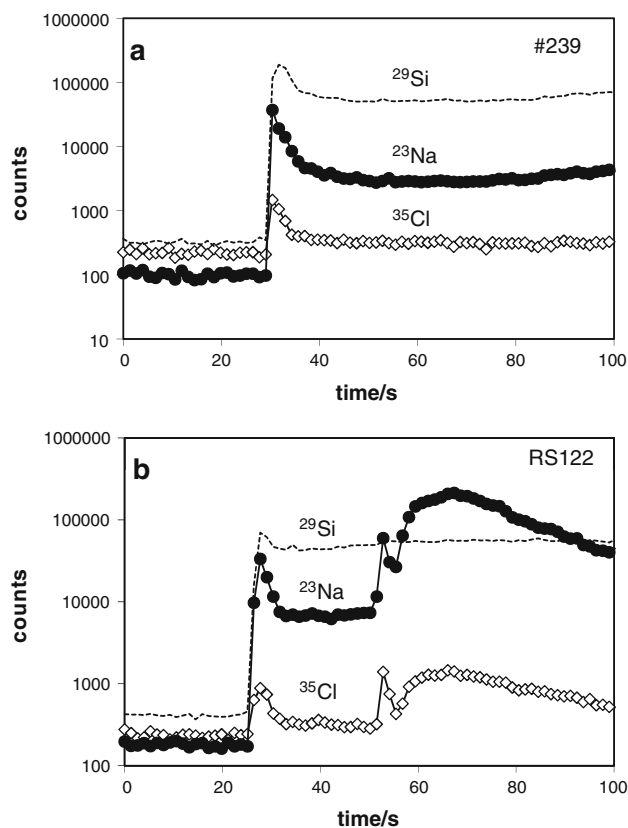


Fig. 2 Time resolved laser ablation-ICP-MS signal for crystals without (a) and with (b) NaCl- or NaCl-rich fluid inclusions. In the latter case only the inclusion-free part (in this case between 30 and 50 s) was used for calculating the Na-content of the sample

In addition, Cl-concentrations are generally too low to allow a significant exchange of OH^- and Cl^- . Thirty-five wt ppm Cl correspond to $100 \text{ Cl}/10^6 \text{ Si}$, which is more than an order of magnitude less than the observed change in $\text{H}/10^6 \text{ Si}$ (Table 1).

From the IR spectra it is evident that all samples host the same type of OH-defects, suggesting that Na^+ does not contribute to the generation of new OH-defect associates (i.e. defects which are not already existing in pure enstatite). In contrast, the negative correlation between the H-content and the Na-content in enstatite (Fig. 3) may be interpreted in a way that Na^+ concurs with H^+ during charge balancing of the Mg-vacancies. This is, however, not very probable, since KCl causes a reduction in H-content similar to NaCl. In RS118 K was below detection limit and therefore (by mole) at least one order of magnitude less abundant than Na in runs with comparable NaCl concentrations in the starting mixture. Therefore the influence on Na^+ on H-incorporation in enstatite can be neglected.

The only remaining important mechanism for the decreased H-incorporation is the reduced water activity due to dissolved salt (Fig. 4). It has to be stressed that—as far

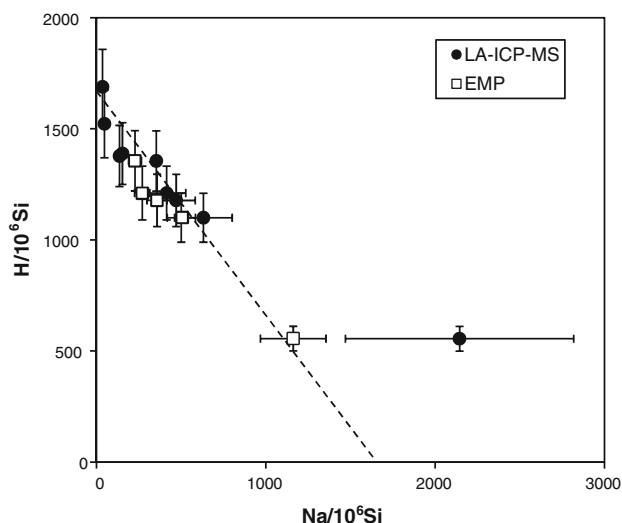


Fig. 3 Hydrogen contents determined by FTIR versus Na contents (determined by laser ablation-ICP-MS and by electron microprobe) of all enstatite crystals

as the nature of the fluid is concerned—there are several unknowns in this study. The solubilities of oxide components in the highly saline fluid are not known and therefore it is not clear, whether aqueous fluid and hydrous melt behave supercritical under the conditions of synthesis. In the case of subcritical behaviour, partition coefficients between aqueous saline fluid and hydrous melt are not known either. The pure system $\text{SiO}_2\text{--H}_2\text{O}$ would behave supercritical (Kennedy et al. 1962) under the final synthesis conditions (i.e. 25 kbar, $1,150^\circ\text{C}$), and approximately 40 wt% silica would be soluble in the fluid (Manning 1994). However, addition of salt at high pressure reduces the silica content of the fluid (Newton and Manning 2000) and may open the miscibility gap between aqueous fluid and hydrous melt. From the results of the experimental work of Newton and Manning (2000) it can be estimated that an increase of X_{NaCl} by 0.1 results in a decrease of silica molality in the fluid of 0.2 log units. Furthermore, as silica is mainly dissolved as dimer ($\text{H}_6\text{Si}_2\text{O}_7$) in high pressure fluid phases (Zotov and Keppler 2002; Newton and Manning 2003), the mole fraction of H_2O would for all experimental runs be only moderately shifted towards lower values. Therefore, only H_2O and the amount of NaCl (or KCl) added to the system was considered for calculating the H_2O mole fraction. The strong initial decline in hydrogen incorporation (Fig. 4) with increasing salinity (decreasing water mole fraction) is independent on the added salt, supporting the above made conclusion that Na-incorporation has only very limited influence on hydrogen incorporation in enstatite (at least in the studied system, where no trivalent cations are present). From H_2O mole fractions, water activities were calculated by two

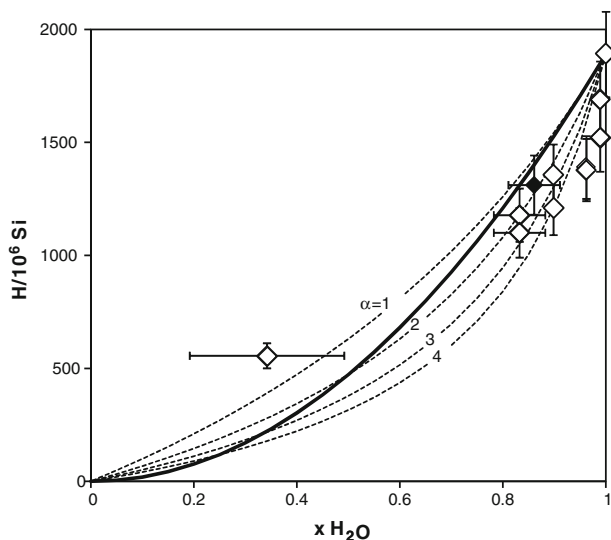


Fig. 4 Influence of the mole fraction of water in the fluid on the hydrogen content in enstatite determined by FTIR. The mole fraction of water was calculated assuming one homogeneous fluid containing the whole inventory of the added salt component. Dissolved silica was not taken into account. The *solid symbol* represents the run conducted with KCl instead of NaCl. The *trend lines* show the expected drop in water activity by the model of Aranovich and Newton (1997) (*solid line*), and the ideal-fused-salt model (Bradley 1962) accounting for varying degree of ionization α of NaCl (*thin dotted lines*) and ideal mixing of Na^+ and Cl^- ions with H_2O molecules. Curves are labelled with the respective value for α

models: (1) H_2O activity is equal to the square of its mole fraction (Aranovich and Newton 1997), and (2) the ideal-fused-salt model involving fully dissociated NaCl and ideal mixing of Na^+ and Cl^- ions with H_2O molecules (Bradley 1962; Newton and Manning 2006), where $a_{\text{H}_2\text{O}} = (1 - X_{\text{NaCl}})/(1 + X_{\text{NaCl}})$. The latter model has been modified taking into account varying degree of ionization α of NaCl, leading to $a_{\text{H}_2\text{O}} = (1 - X_{\text{NaCl}})/(1 + \alpha X_{\text{NaCl}})$ (Aranovich and Newton 1996), where the degree of ionization depends on the density of water under run conditions. $\rho_{\text{H}_2\text{O}}$ was obtained by extrapolation from Halbach and Chatterjee (1982). For 25 kbar and 1,150°C we obtain $\rho_{\text{H}_2\text{O}} = 1.075 \text{ g/m}^3$, leading to $\alpha = 1.46$. Using this parameter we obtain—except for the most water-rich samples—a fair accordance to the experimental data points (Fig. 4). The mismatch of the most water-poor data point may be due to its location above the salt saturation level (salt inclusions have been detected during laser ablation analysis, Fig. 2), and its actual H_2O mole fraction of may be higher. It may also be possible that α changes with mole fraction from initially high values at pure water towards smaller values at higher salt concentrations, which would both explain the sudden decrease of water in enstatite at low NaCl mole fractions and the positive deviation of the data point at high NaCl mole fraction.

Element partitioning

It is not known, whether at synthesis conditions a hydrous melt coexists with an aqueous fluid or whether one homogeneous fluid exists. Therefore, there are several ways to estimate the partitioning behaviour of Na and H. A lower estimate for the partition coefficient $D_{\text{Na}}^{\text{En/FI}}$ can be obtained, if it is considered that a highly saline aqueous fluid, which contains most of the Na and Cl in the system, is coexisting with a silicate melt. The partitioning of Na between enstatite and fluid can then be calculated, considering the silica content dissolved in the aqueous fluid at 25 kbar and 1,150°C (Table 1) based on the experimental results of Manning (1994) and Newton and Manning (2000). The obtained value $D_{\text{Na}}^{\text{En/FI}} = 0.0013$ (Fig. 5) is significantly higher than $D_{\text{H}}^{\text{En/FI}} = 0.0003$ (1889 H/10⁶Si corresponding to 170 wt ppm H_2O in enstatite, divided by 60 wt% H_2O in the fluid). An upper estimate for element partitioning can be performed, if the bulk starting composition of each run is compared to Na and H content of the synthesized crystals. In this case slightly higher partition coefficients (i.e. $D_{\text{Na}}^{\text{En/Startmix}} = 0.0027$, $D_{\text{H}}^{\text{En/Startmix}} = 0.0006$) are obtained (Fig. 5), but the ratio of the two coefficients stays the same. This means that Na is much more compatible in enstatite than H, which substantiates the above made conclusion that—at its best—only a small fraction of the incorporated Na^+ -ions actually compensate OH-defects in the enstatite lattice and by far most of the Na^+ has to be incorporated by another substitution

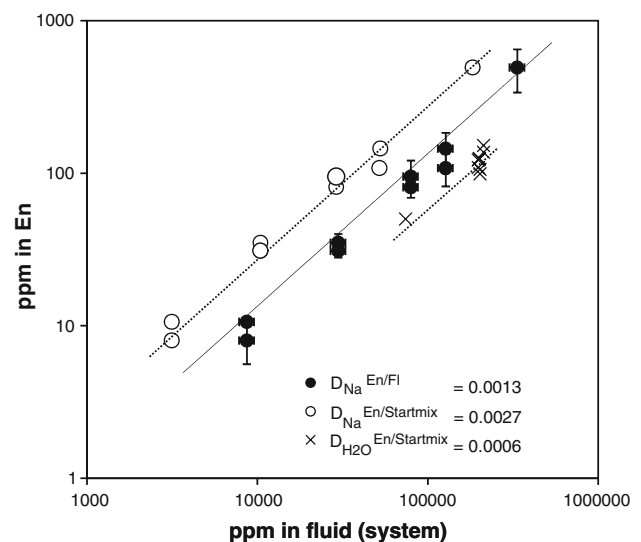


Fig. 5 Na partitioning and H_2O partitioning between enstatite and fluid. Na concentrations in enstatite were determined by laser ablation-ICP-MS. Na concentrations in the fluid were calculated assuming one homogeneous fluid containing the whole inventory of the added salt component and a dissolved silica component (Table 1) estimated from the studies of Manning (1994) and Newton and Manning (2000)

mechanism. It can, however, be excluded that Na is incorporated as NaCl or NaCl-fluid inclusions, since there is no apparent correlation between the Na- and the Cl-content of the enstatite crystals.

An upper limit for Cl-partitioning can also be estimated from the results of the most Cl-rich run RS122. Considering electron microprobe results (Table 1) as upper limit for the real Cl incorporated into the enstatite lattice, $D_{\text{Cl}}^{\text{En/FI}}$ is at least one order of magnitude lower than $D_{\text{Na}}^{\text{En/FI}}$. This means that Cl is the most incompatible volatile in our study, in good agreement to estimates based on natural systems (Saal et al. 2002; Straub and Layne 2003).

Implications for the Earth's mantle

Estimates for the total concentration of chlorine in the upper mantle range from 1 ppm for depleted MORB mantle (Saal et al. 2002) to 17 ppm for pyrolite (McDonough and Sun 1995), where a large proportion of the subducted chlorine is transferred to the mantle reservoir (Philippot et al. 1998). Models for the global chlorine cycle suggest a balance between input through subduction and output through magmatism (Ito et al. 1983; Straub and Layne 2003). With respect to chlorine, the results of our study are somewhat speculative. However, the most salt-rich experimental run of our experiments (RS122), where solid NaCl-inclusions have been inferred, sets an upper estimate (i.e. 30 ppm) for Cl-incorporation in enstatite at 25 kbar and 1,150°C. This value is similar to the Cl-content of an orthopyroxene from a high pressure metamorphic (>700°C, 18 kbar) harzburgite converted from an antigorite serpentinite during subduction metamorphism (Scambelluri et al. 2004), where 25 ppm Cl were detected. Considering the limited Cl incorporation in our experiments even at salt saturation, orthopyroxene does not seem to be a major host for Cl in the mantle, unless Cl is incorporated as fluid inclusions (Philippot et al. 1998).

On the other hand, the results of our study imply that H incorporation in nominally anhydrous minerals in NaCl-rich portions of the mantle (e.g. subduction settings) may be more limited than previously thought. Low water contents in pyroxenes from the sub-arc mantle wedge have previously been interpreted to be controlled by the redox state of the mantle rock (Peslier et al. 2002), since a negative correlation between oxygen fugacity and water content of the pyroxenes could be established. On the basis of our experimental results the salinity of the fluid phase could be an alternative explanation. This aspect has so far not been investigated in natural samples and deserves further scrutiny. Furthermore, it has to be taken into consideration that water incorporation in orthopyroxene is strongly enhanced by the presence of trivalent cations, such as Al^{3+} (Stalder, 2004; Mierdel et al. 2007). For natural orthopyroxenes, in

which these cations usually reach high concentrations (e.g. typically up to several wt% Al_2O_3), the presence of monovalent cations, such as Na^+ , may have a negative effect on the water storage capacity.

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